






# different directions for electron pairs	Electron-pair geometry	Hybridization	# lone pairs	Molecular shape	Examples
2	linear 	sp	0	linear	O=C=O Cl-Be-Cl H-C≡N H-C≡C-H
3	trigonal planar 	sp ²	0	trigonal planar	BF ₃ CO ₃ ²⁻
			1	bent	GeF ₂ , NO ₂ ⁻
4	tetrahedral 	sp ³	0	tetrahedral	CH ₄
			1	trigonal pyramidal	NH ₃
			2	bent	HOH
5	trigonal bipyramidal 	dsp ³	0	trigonal bipyramidal	PCl ₅
			1	seesaw	SF ₄
			2	T-shaped	ClF ₃
			3	linear	XeF ₂
6	octahedral 	d ² sp ³	0	octahedral	SF ₆
			1	square pyramidal	IF ₅
			2	square planar	XeF ₄